

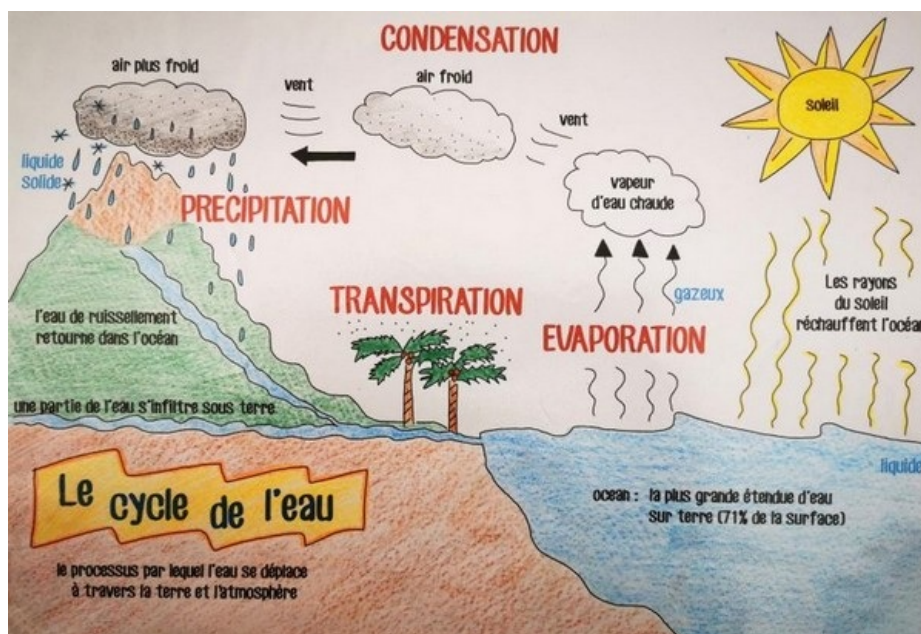
Sommaire

- 1- L'eau dans notre environnement
 - 2- Les états physiques de la matière
 - 3- Mesure du volume des solides et des liquides
 - 4- Mesure de la masse des solides et des liquides
 - 5- La masse volumique
 - 6- La pression et la pression atmosphérique
 - 7- Le modèle particulaire de la matière
 - 8- La chaleur et les changements d'état physique de la matière
 - 9- Les mélanges
 - 10- La dissolution dans l'eau
 - 11- La séparation des constituants d'un mélange
 - 12- Le corps pur et ses caractéristiques
 - 13- Le traitement des eaux
-

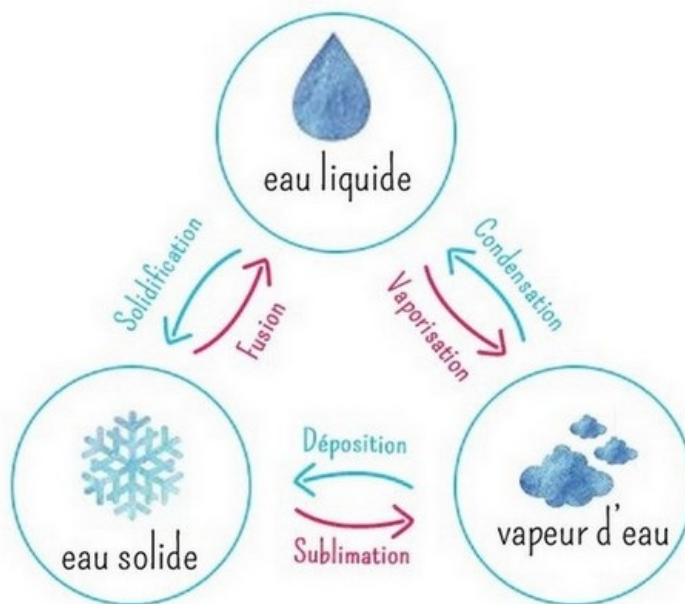
- 1- L'eau dans notre environnement



Mers et océans	97,2%
Glaces polaires et glaciers	2,15%
Eaux souterraines	0,63%
Lacs et rivières	0,01%
Atmosphère	0,001%

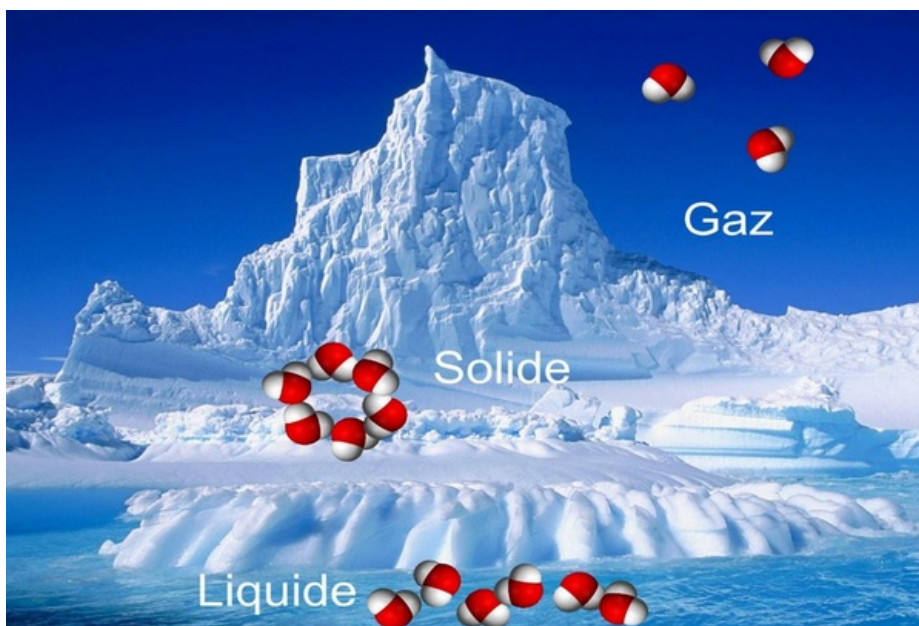


LES DIFFÉRENTS ÉTATS DE L'EAU





2- Les états physiques de la matière



	Les solides compacts	Les solide non compacts (divisés)
Les propriétés physique	<ul style="list-style-type: none"> • Faits d'un seul bloc • Ont une forme propre • Ont un volume constant • On peut les saisir avec les doigts 	<ul style="list-style-type: none"> • Formés de grains ou fragment très petits • N'ont pas une forme propre (ils prennent la forme de récipient dans lequel ils sont placés) • On ne peut pas les saisir avec les doigts
Exemples	Un bol, un cahier, un crayon	La farine, le café moulu, le sel

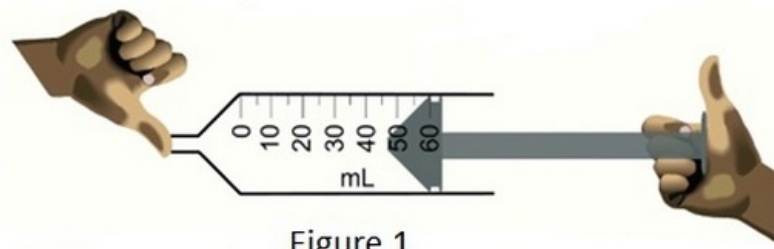


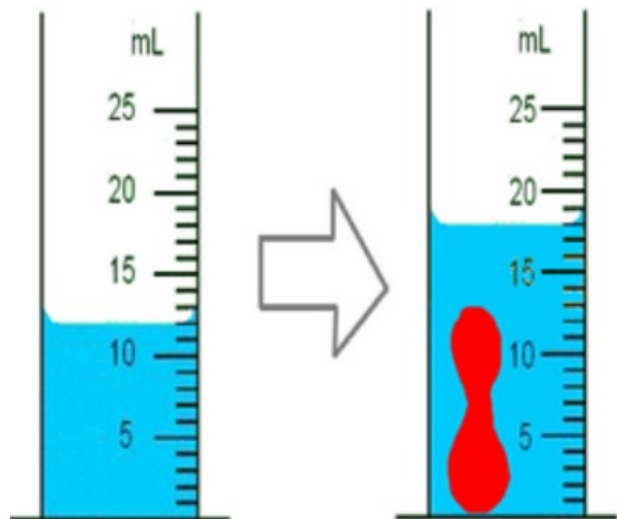
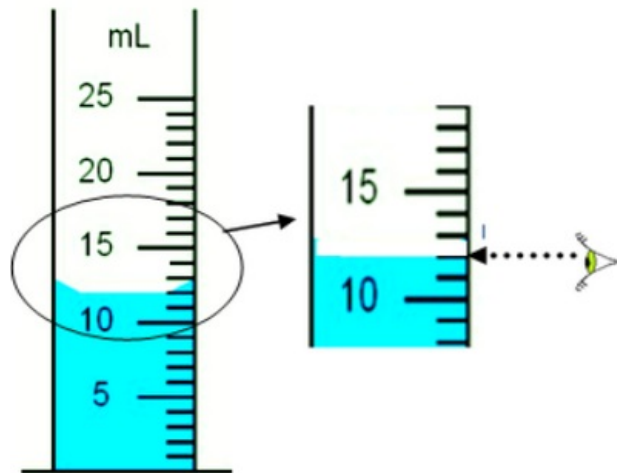
Figure 1

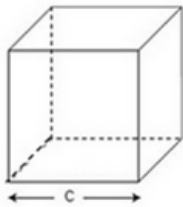

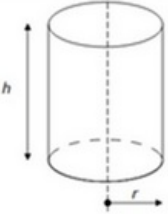
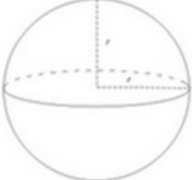


3- Mesure du volume des solides et des liquides



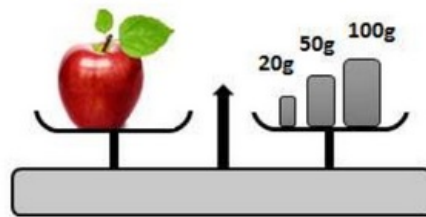
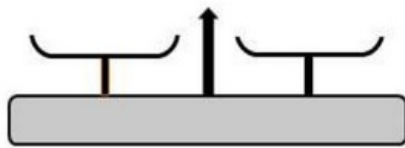
hm ³			dam ³			m ³			dm ³			cm ³			mm ³		
										L		dl	cl	ml			



solide	cube	Parallépipède rectangle	cylindre	sphère
Forme géométrique				
Volume V	$V = c \times c \times c$ $V = c^3$	$V = L \times l \times h$	$V = \pi \times r \times r \times h$ $V = \pi \times r^2 \times h$	$V = \frac{4}{3} \times \pi \times r \times r \times r$ $V = \frac{4}{3} \times \pi \times r^3$

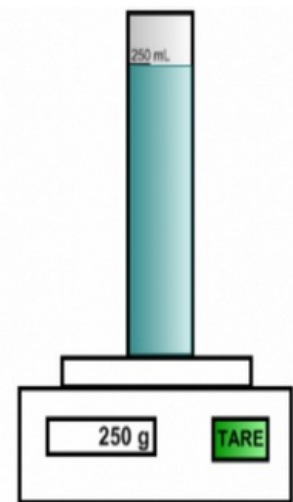
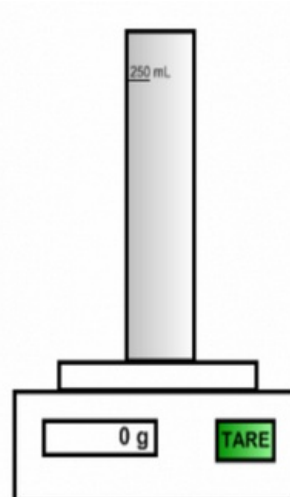
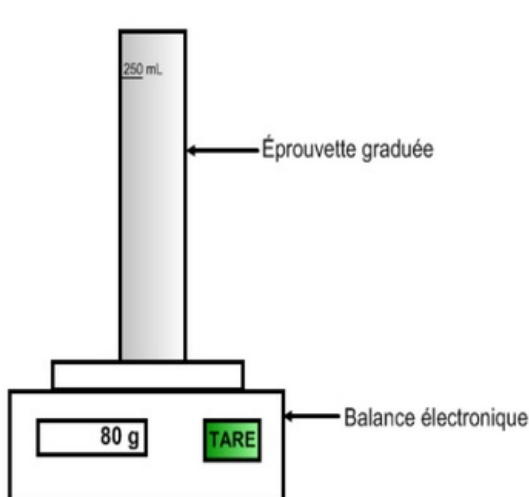
4- Mesure de la masse des solides et des liquides

Unité	tonne	quintal	kilogramme	hectogramme	décagramme	gramme	décigramme	centigramme	milligramme
Symbole	t	q	kg	hg	dag	g	dg	cg	mg

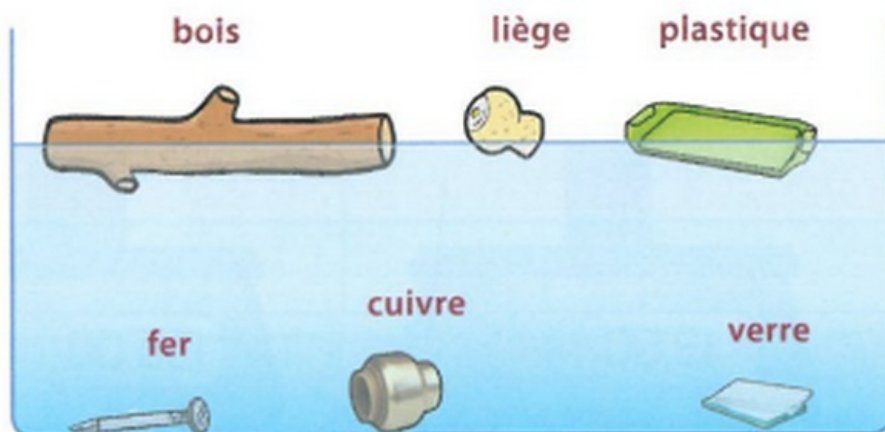
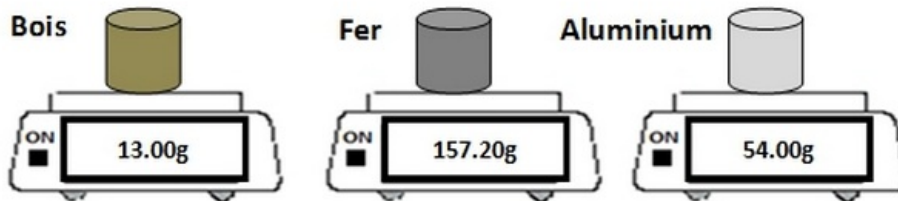
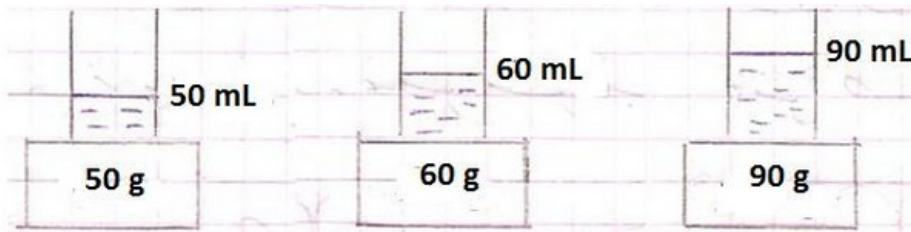
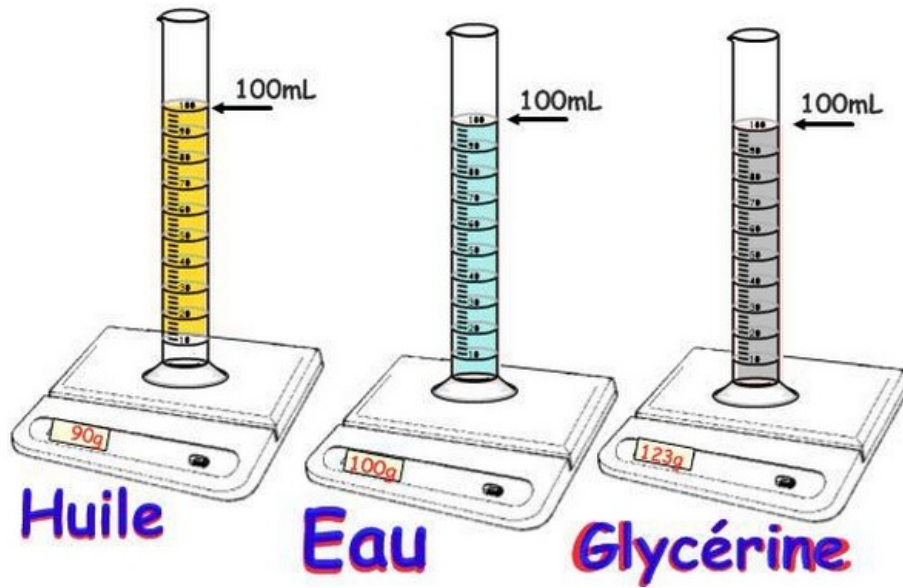


la balance est remise à zéro

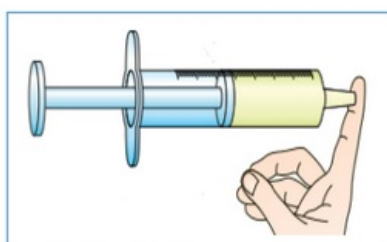
m = 280 g



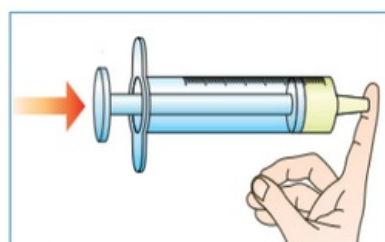
5- La masse volumique



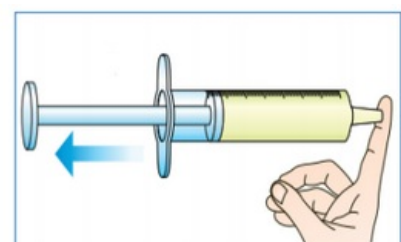
6- La pression et la pression atmosphérique



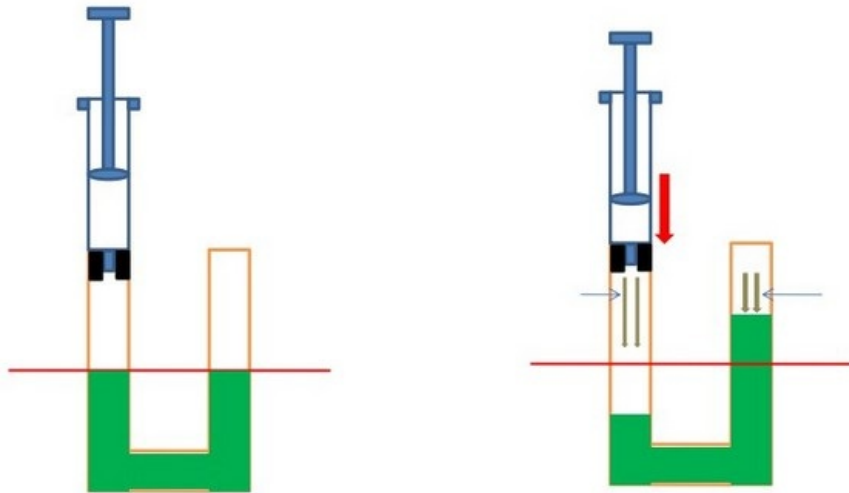
On bouche la seringue



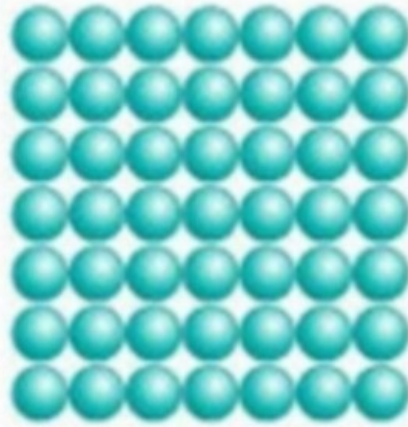
Cas N° 1 : On pousse le piston



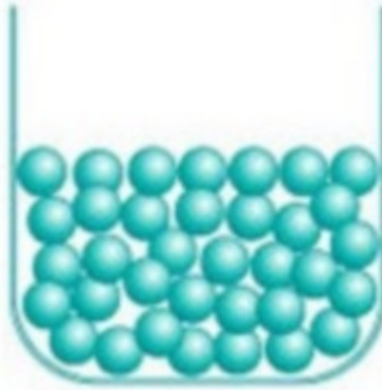
Cas N° 2 : On tire le piston



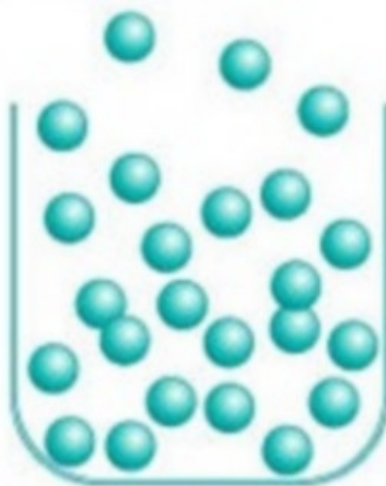
7- Le modèle particulaire de la matière



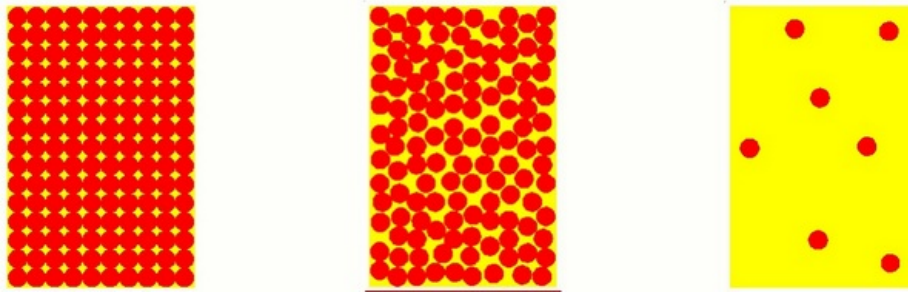
solide



liquide

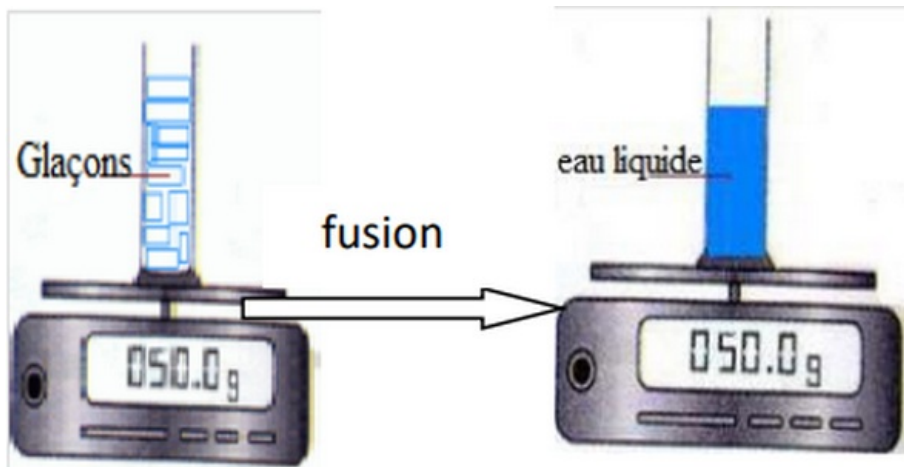
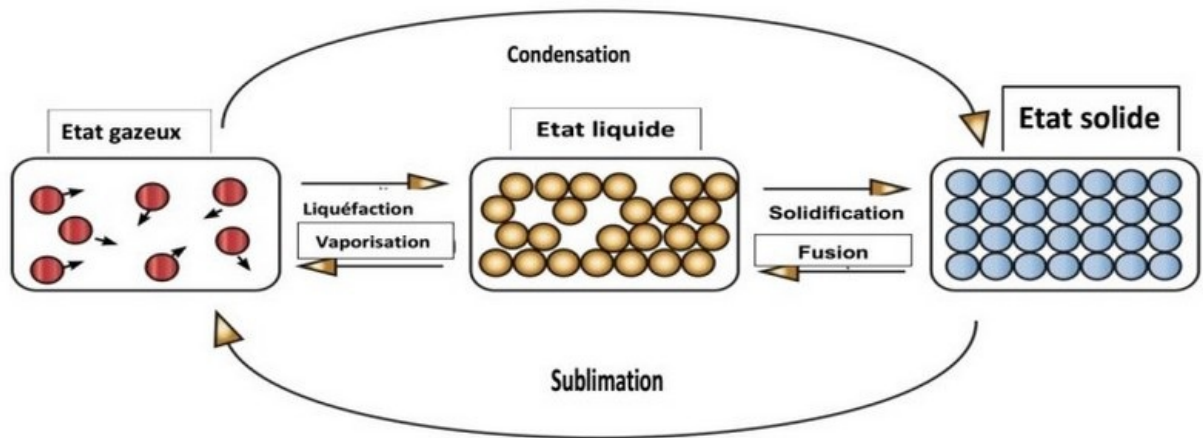
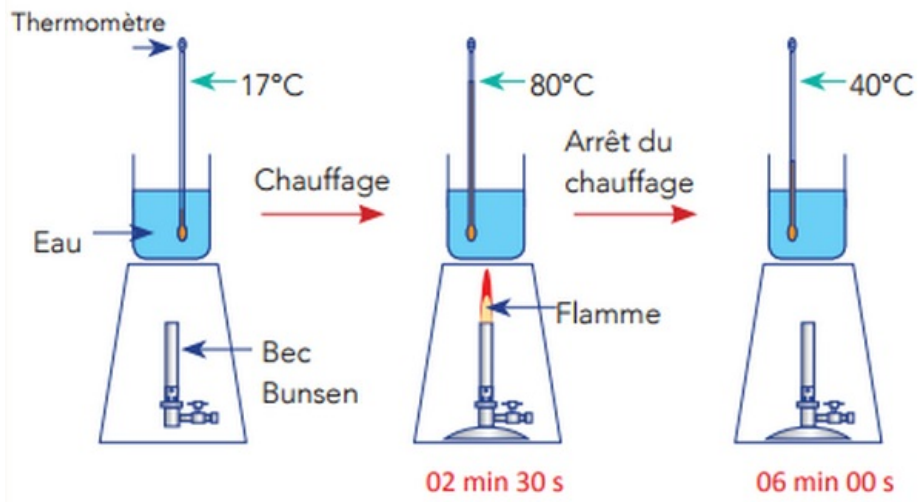


gaz



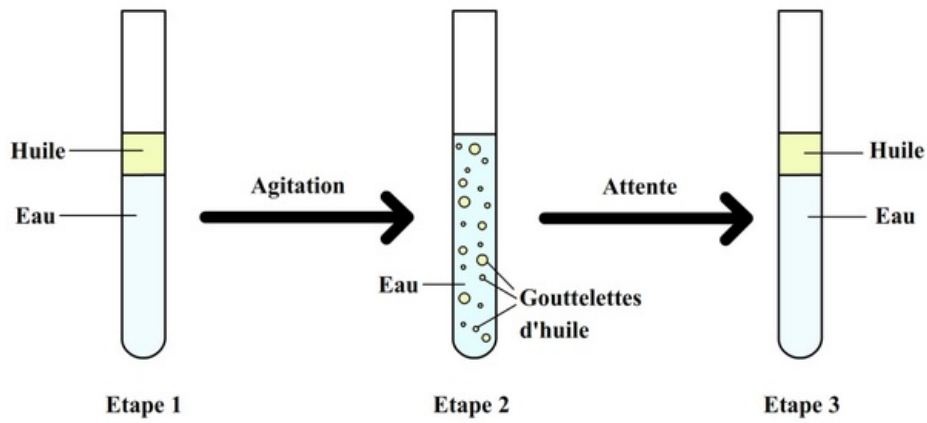
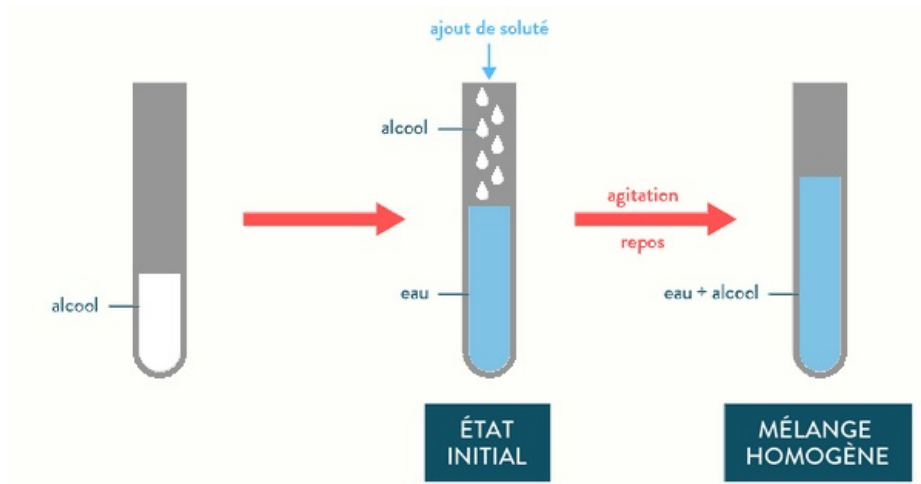
8- La chaleur et les changements d'état physique de la matière



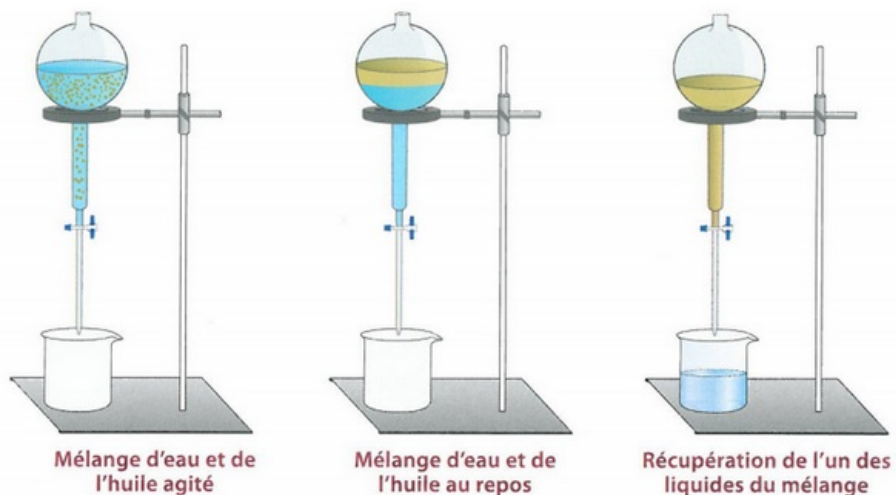
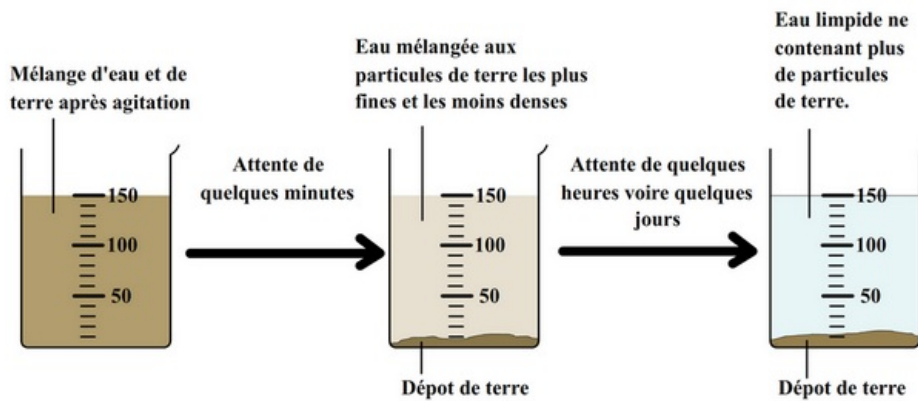


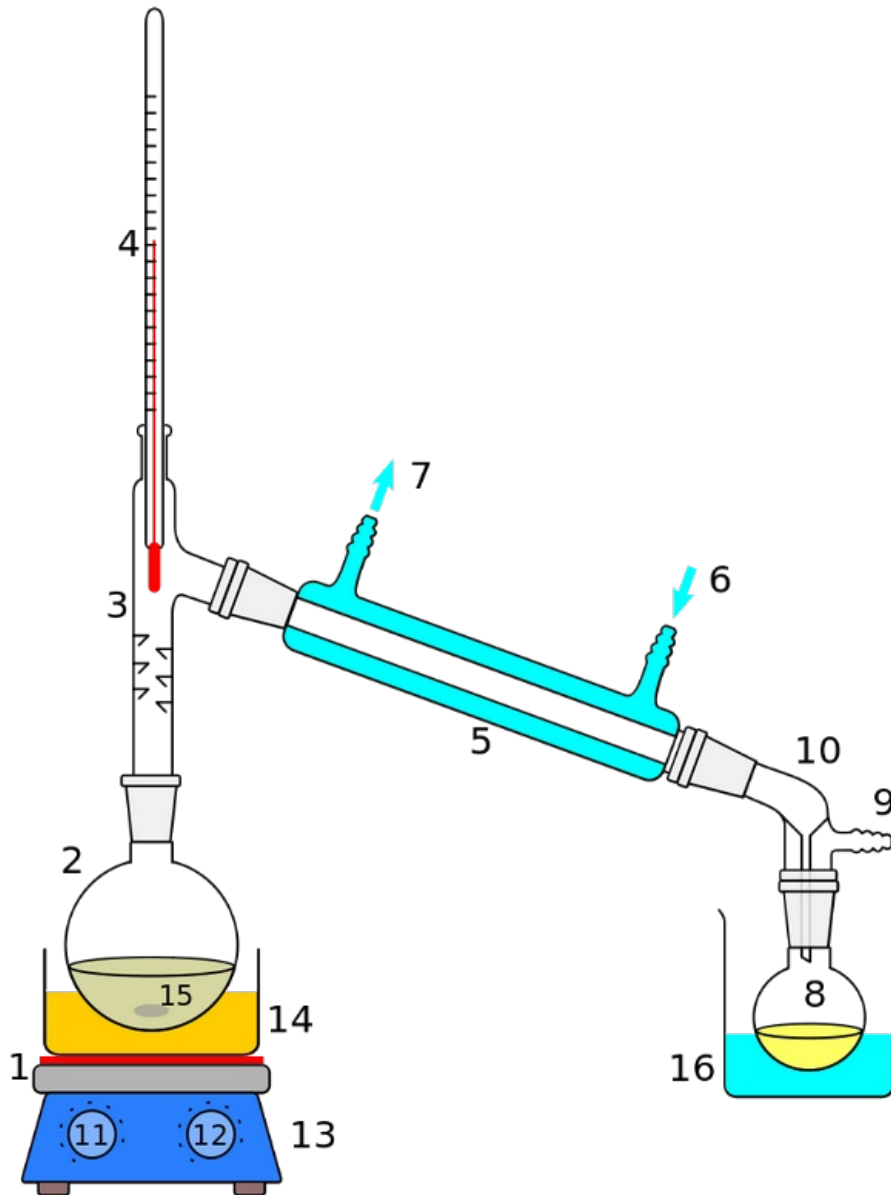
9- Les mélanges



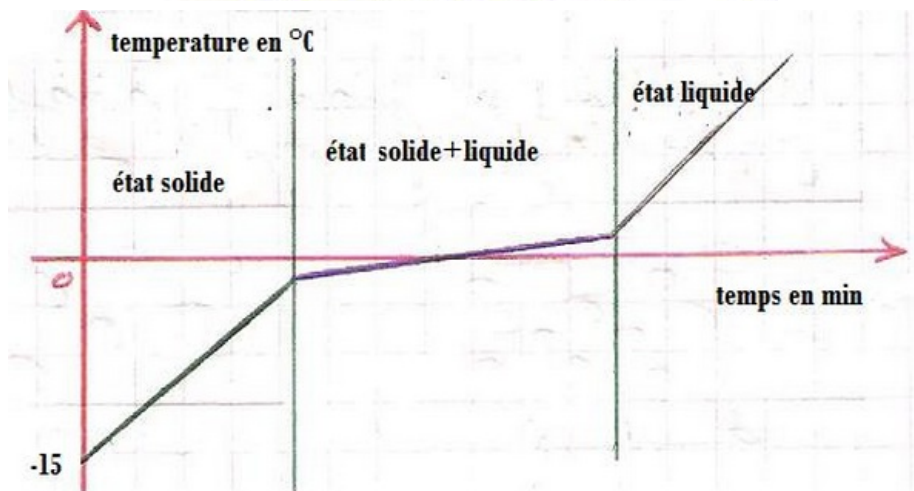
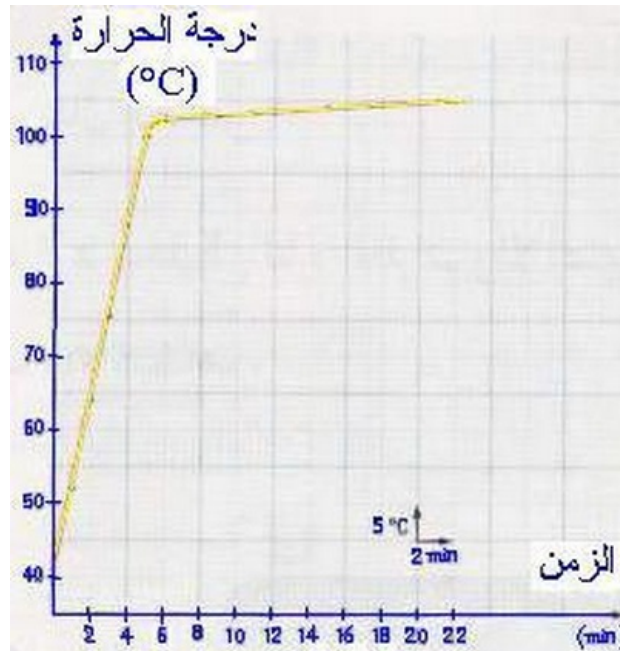
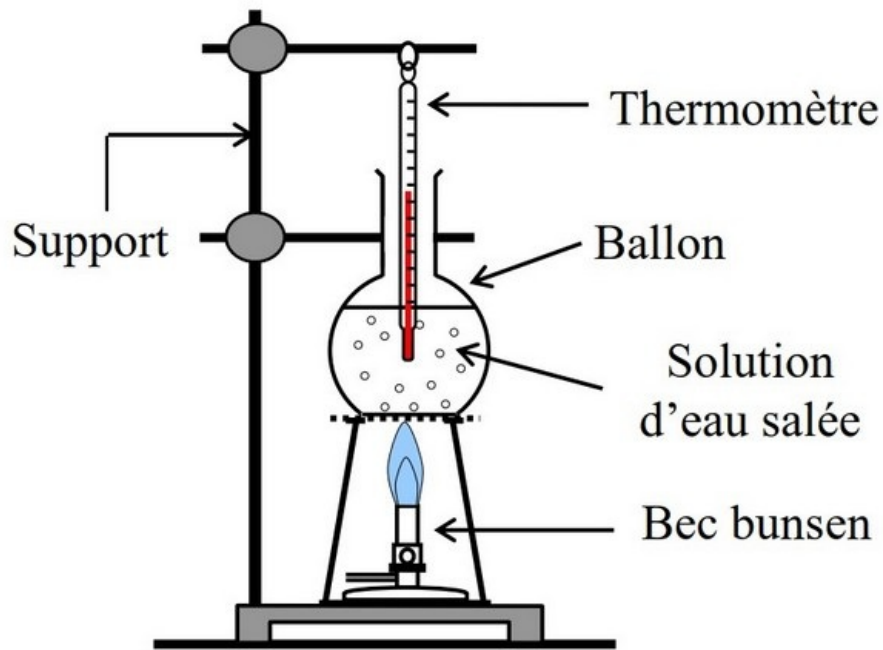


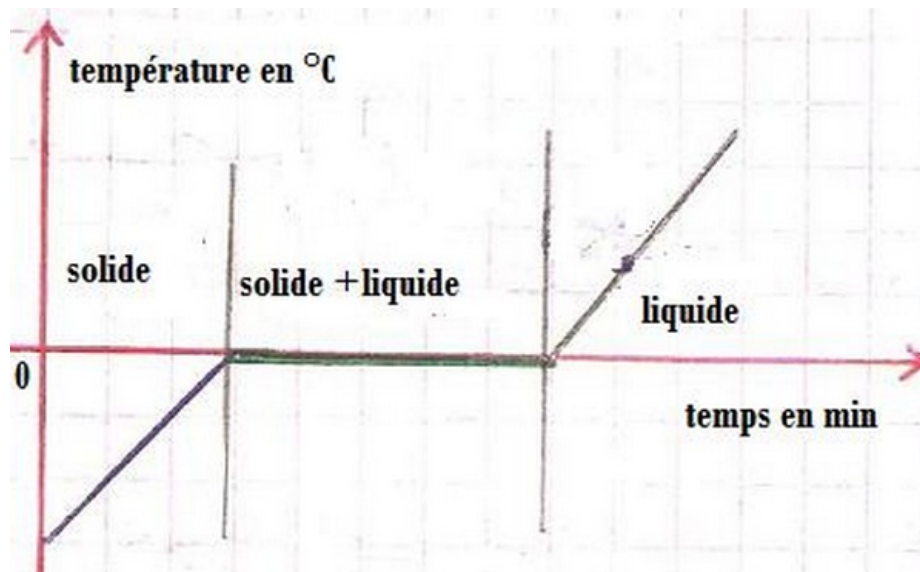
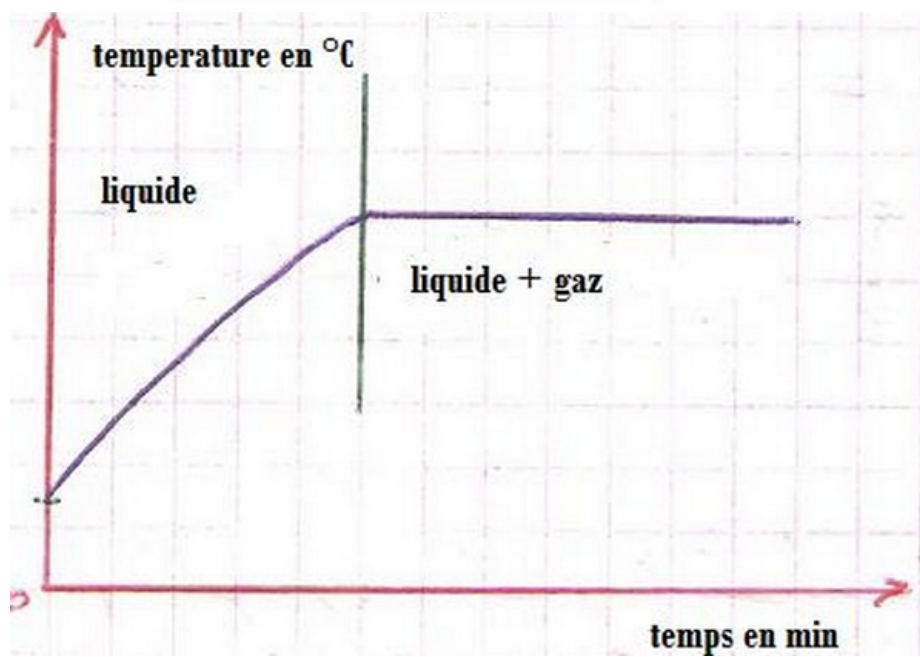
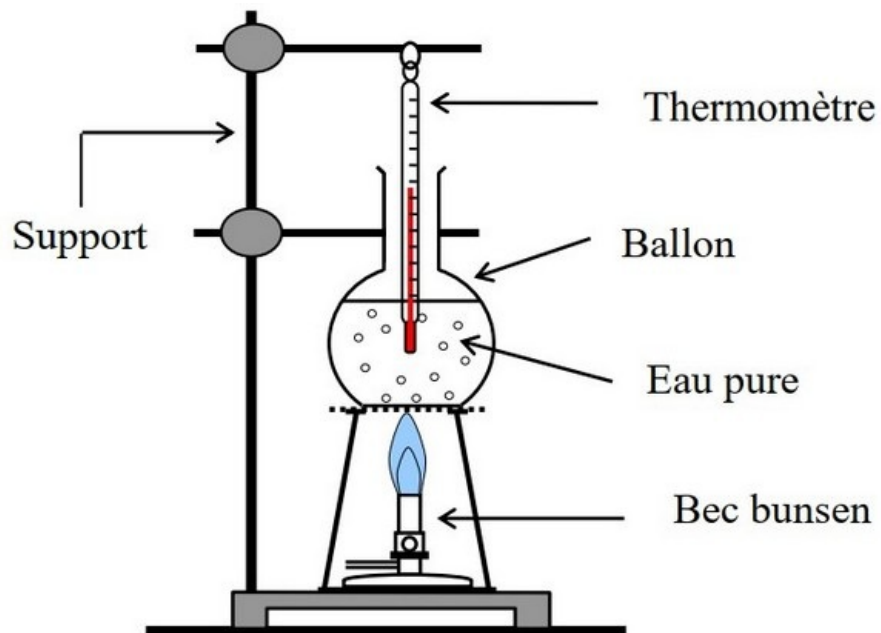
10- La dissolution dans l'eau



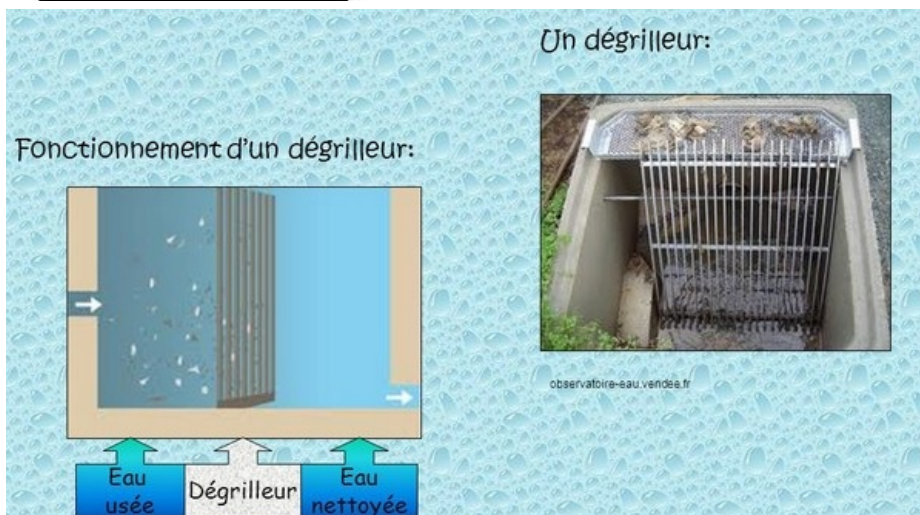
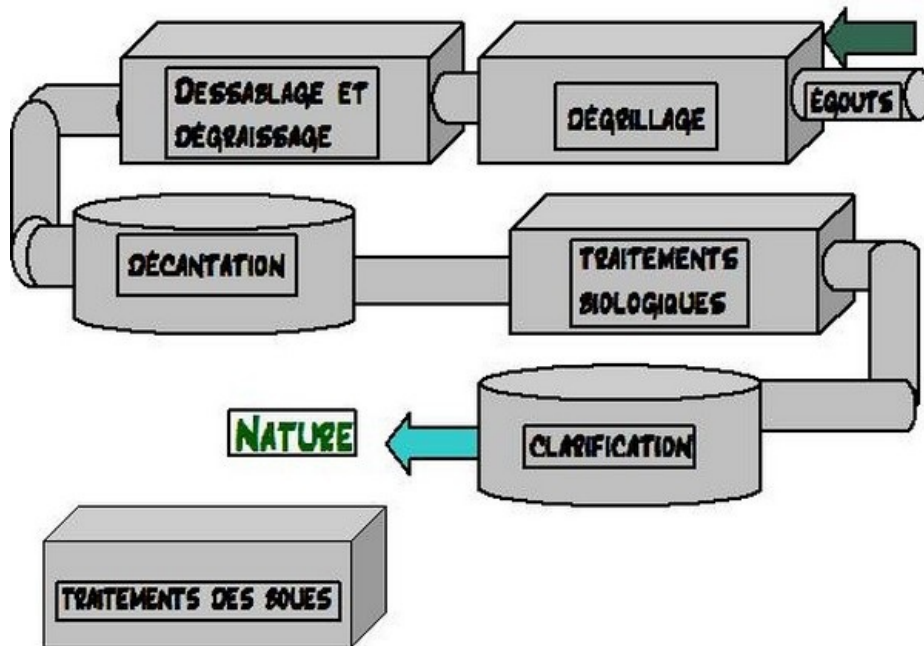
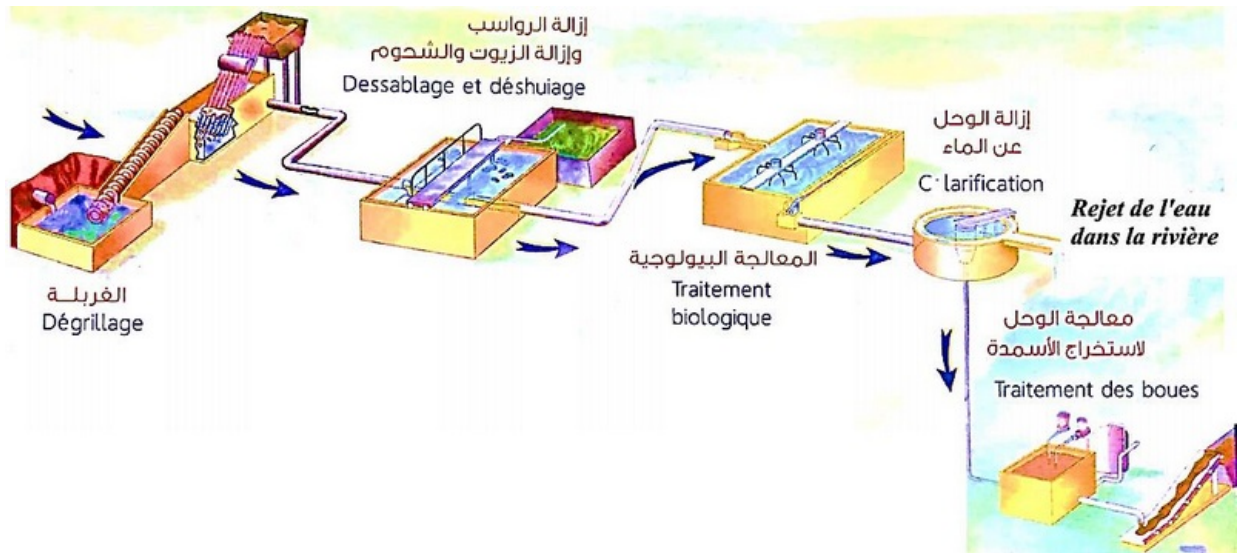


12- Le corps pur et ses caractéristiques





13- Le traitement des eaux



Un dessableur:

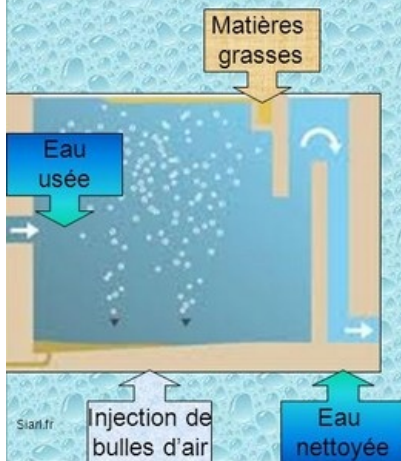
Pour le moment, nous nous intéressons seulement à la partie dessablage:



step.ouvaton.org

Un dégraisseur:

Maintenant, nous nous intéressons à la partie dégraisage:



ecoledeleau.eau-artois-picardie.fr

